

4. 65.0 kilojoules of energy are added to 150 g of ice at 0.0°C . What is the final temperature of the water?
5. 250 kJ of energy are removed from a 4.00×10^2 g sample of water at 60°C . Will the sample of water completely freeze? Explain.
6. An ice cube tray full of ice (235g) at -7.0°C is allowed to warm up to room temperature (22°C). How much energy must be absorbed by the contents of the tray in order for this to happen?
7. If this same quantity of energy were removed from 40.0 g of water vapor at 100°C , what would be the final temperature of the water?