

Name \_\_\_\_\_

## Ionic Compounds

	Name	Formula	Cation	Anion
1.	Potassium Nitride	$K_3N$	$K^+$	$N^{3-}$
2.	Lead (II) Sulfide	$PbS$	$Pb^{2+}$	$S^{2-}$
3.	Ammonium Oxide	$(NH_4)_2O$	$NH_4^+$	$O^{2-}$
4.	Magnesium Nitrate	$Mg(NO_3)_2$	$Mg^{2+}$	$NO_3^-$
5.	Aluminum Bromide	$AlBr_3$	$Al^{3+}$	$Br^-$
6.	Calcium Phosphide	$Ca_3P_2$	$Ca^{2+}$	$P^{3-}$
7.	Zinc Oxide	$ZnO$	$Zn^{2+}$	$O^{2-}$
8.	Sodium Iodide	$NaI$	$Na^+$	$I^-$
9.	Barium Fluoride	$BaF_2$	$Ba^{2+}$	$F^-$
10.	Copper (I) Sulfide	$Cu_2S$	$Cu^+$	$S^{2-}$
11.	Sodium Phosphate	$Na_3PO_4$	$Na^+$	$PO_4^{3-}$
12.	Iron (II) Acetate	$Fe(C_2H_3O_2)_2$	$Fe^{2+}$	$C_2H_3O_2^-$
13.	Aluminum Sulfate	$Al_2(SO_4)_3$	$Al^{3+}$	$SO_4^{2-}$
14.	Copper (II) Oxide	$CuO$	$Cu^{2+}$	$O^{2-}$
15.	Cobalt (II) Hydroxide	$Co(OH)_2$	$Co^{2+}$	$OH^-$
16.	Ammonium BiCarbonate	$NH_4HCO_3$	$NH_4^+$	$HCO_3^-$
17.	Calcium Nitride	$Ca_3N_2$	$Ca^{2+}$	$N^{3-}$
18.	Iron (III) Bromide	$FeBr_3$	$Fe^{3+}$	$Br^-$
19.	Silver Hypochlorite	$AgClO$	$Ag^+$	$ClO^-$
20.	Nickel (II) Phosphide	$Ni_3P_2$	$Ni^{2+}$	$P^{3-}$
21.	Cobalt (III) Carbonate	$Co_2(CO_3)_3$	$Co^{3+}$	$CO_3^{2-}$
22.	Ammonium Sulfate	$(NH_4)_2SO_4$	$NH_4^+$	$SO_4^{2-}$
23.	Zinc Chloride	$ZnCl_2$	$Zn^{2+}$	$Cl^-$
24.	Lithium Hydroxide	$LiOH$	$Li^+$	$OH^-$
25.	Aluminum Nitrate	$Al(NO_3)_3$	$Al^{3+}$	$NO_3^-$