

Name _____

Protons, neutrons, and electrons

Symbol	Name	At #	At Mass	p ⁺	n ^o	e ⁻	Charge	Atom or Ion
$^{31}_{15}\text{P}^0$	Phosphorus-31	15	31	15	16	15	0	atom
$^{27}_{13}\text{Al}^0$	Aluminum-27	13	27	13	14	13	0	atom
$^{67}_{30}\text{Zn}^0$	Zinc-67	30	67	30	37	30	0	atom
$^{56}_{26}\text{Fe}^{3+}$	Iron-56	26	56	26	30	23	3+	ion
$^{88}_{38}\text{Sr}^0$	Strontium-88	38	88	38	50	38	0	atom
$^{30}_{14}\text{Si}^0$	Silicon-30	14	30	14	16	14	0	atom
$^{39}_{19}\text{K}^{1+}$	Potassium-39	19	39	19	20	18	1+	ion
$^{23}_{11}\text{Na}^0$	Sodium-23	11	23	11	12	11	0	atom
$^{112}_{48}\text{Cd}^0$	Cadmium-112	48	112	48	64	48	0	atom
$^{91}_{40}\text{X}^{2-}$	Ex-91	40	91	40	51	42	2-	ion

Symbol	Name	At #	At Mass	p ⁺	n ^o	e ⁻	Charge	Atom or Ion
$^{23}_{11}\text{Na}^{1+}$	Sodium-23	11	23	11	12	10	1+	ion
$^{19}_{9}\text{F}^{1-}$	Fluorine-19	9	19	9	10	10	1-	ion
$^{208}_{82}\text{Pb}^{4+}$	Lead-208	82	208	82	126	78	4+	ion
$^{109}_{47}\text{Ag}^{1+}$	Silver-109	47	109	47	62	46	1+	ion
$^7_3\text{Li}^0$	Lithium-7	3	7	3	4	3	0	atom
$^{32}_{16}\text{S}^{2-}$	Sulfur-32	16	32	16	16	18	2-	ion
$^{138}_{56}\text{Ba}^{2+}$	Barium-138	56	138	56	82	54	2+	ion
$^{14}_7\text{N}^{3-}$	Nitrogen-14	7	14	7	7	10	3-	ion
$^{86}_{38}\text{Sr}^0$	Strontium-86	38	86	38	48	38	0	atom
$^{58}_{28}\text{Ni}^{2+}$	Nickel-58	28	58	28	30	26	2+	ion